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Stoichiometry Section 11.1

What is stoichiometry? 1.

true 2. true 3. false 4.

true 5. true 6. 2, 2, 64.10

7. 3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

methanol and oxygen gas 11.

carbon dioxide and water 12.

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160.10 g 13. 160.10 g 14.

They are equal. 15. A mole ratio is a ratio between the numbers of moles

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Stoichiometry Section

11.1.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

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Class Stoichiometry Section

11.1 What is stoichiometry?

In your textbook, read about stoichiometry and the

balanced equation. For each statement below, write true or false.

2. 3. 4. 5. The

study of the quantitative relationships between the

amounts of reactants used

and the amounts of products

formed by a chemical

reaction

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• stoichiometry • mole ratio
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Number of Moles

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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